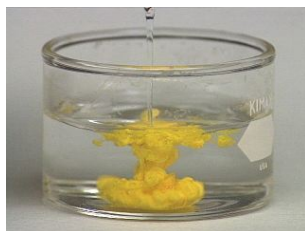


Demonstration Download: Precipitation in Chemical Changes

ChemTeacher Demonstrations

Precipitation in Chemical Changes

Author:



Adding very pale yellow sodium iodide to colorless lead(II) nitrate produces a yellow precipitate.

□

Futher Images:



Materials

Solution A: 0.5 M sodium iodide

Solution B: 0.2 M lead(II) nitrate

Procedure

Adding sodium iodide to colorless lead(II) nitrate.

Notes

Solution A: 0.5 M sodium iodide (very pale yellow)

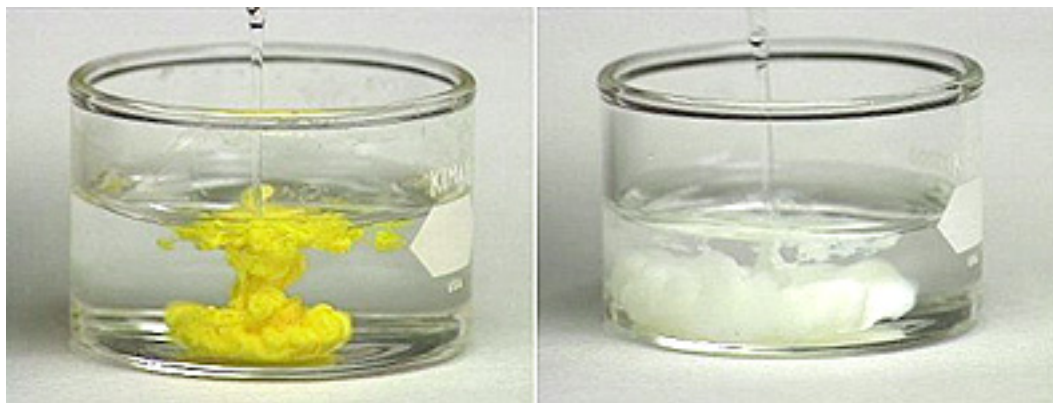
Solution B: 0.2 M lead(II) nitrate (colorless)

Precipitate: (yellow)

$\text{Pb}(\text{NO}_3)_2(\text{aq}) + 2 \text{NaI}(\text{aq}) \rightarrow \text{PbI}_2(\text{s}) + 2 \text{NaNO}_3(\text{aq})$

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ChemTeacher Demonstrations



Demonstration Time	minutes	Difficulty	Some laboratory experience required
Portions	Not specified	Availability of Materials	
Preparation Time		Cost of materials	